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~~?G = 0 at equilibrium so K = 1 in ?G = -RT ln K. (In the absence of these, ?G = -RT ln K gets 1 point). The above concludes the AP scoring standards published in 1991. The following is simply alternate. ways of answering which the AP readers may or may not have given full credit to.~~

~~AP* Bonding & Molecular Structure Free Response Questions ...~~

~~KEY AP* Bonding & Molecular Structure Free Response Questions page I Essay Questions 1991 a) two points AS will be negative. The system becomes more ordered as two gases form a solid. b) two points AH must be negative. For the reaction to be spontaneous, AG must be negative, so AH must be more negative than -TAS is positive.~~

~~KM 554e-20161102150929~~

~~Advanced Placement Chemistry: 1991 Free Response Questions 1) The acid ionization constant, Ka, for propanoic acid, C₂H₅COOH, is 1.3 x 10⁻⁵. (a) Calculate the hydrogen ion concentration, [H⁺], in a 0.20-molar solution of propanoic acid. [VIEW ANSWER] [Find Similar] ap10 chemistry scoring guidelines final~~

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~~Old AP Chemistry Exams Pre 1995~~

Advanced Placement Chemistry: 1990 Free Response Answers [Δ] and [Σ] are used to indicate the capital Greek letters. [$\sqrt{\quad}$] applies to the numbers enclosed in parenthesis immediately following

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~~John A. Ferguson Senior High School~~

Kinetics Free Response Practice 5) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightarrow 2 \text{HI}(\text{g})$ For the exothermic reaction represented above, carried out at 298 K, the rate law is as follows. $\text{Rate} = k [\text{H}_2] [\text{I}_2]$ Predict the effect of each of the following changes on the initial rate of the reaction and explain your prediction.

~~Kinetics Free Response Practice~~

$P = \frac{nRT}{V} = \frac{(0.02976 \text{ mol})(0.0821 \text{ L atm mol}^{-1}\text{K}^{-1})(483 \text{ K})}{1.18 \text{ atm}}$ Note: The first point is earned for correctly setting up either of these expressions. The second point is earned for the correct answer (or an answer consistent with the numbers used). Also, combining $P_{\text{total}} V = n_{\text{total}}RT$ with $P_{\text{total}} = \frac{n_{\text{total}}RT}{V}$

~~2000 AP Chemistry Scoring Guidelines—College Board~~

Free Response Question (1991) 2) The molecular formula of a hydrocarbon is to be determined by analyzing its combustion products and investigating its colligative properties. (a) The hydrocarbon burns completely, producing 7.2 grams of water and 7.2 liters of CO_2 at standard conditions. Find the empirical formula.

~~AP Chemistry study guide for Solutions (Chapter 11)~~

2020 Free-Response Questions. 2020 free-response questions are now in the AP Classroom question bank for teachers to assign to students as homework or in class, and do not require secure assessment. These questions have been updated, where possible, to best match the format of free-response questions in the course and exam description and on traditional AP Exams.

~~AP Chemistry Exam—AP Central | College Board~~

AP* Thermodynamics Free Response Questions Essay Questions 1991 $\text{BC}_{13}(\text{g}) + \text{C}_{13}\text{BNH}_3(\text{s})$ The reaction represented above is a reversible reaction. page 1 (a) Predict the sign of the entropy change, ΔS , as the reaction proceeds to the right, Explain your prediction.

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