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Reduction Questions 1.

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Explain in your own words the following oxidation processes: a.

Gain of oxygen b. Loss of hydrogen c. Loss of electrons 2. Explain in

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your own words the following reduction processes: a. Loss of oxygen b.

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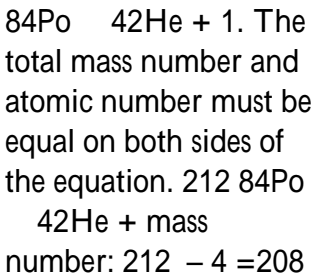
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704 CHAPTER 22

Identify the product that
balances the following
nuclear reaction:



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atomic number: 84 – 2

=82 2. The nuclide has
a mass number of 208
and an atomic number

of 82, 208 82Pb. 3.

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Chapter 22-23

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22.2 Structures of

Alkenes & Alkynes SG

22.3 & 22.4 SG 23.1,

23.2, 23.3 Chapter 22

Supplemental Problems

Structures of Functional

Groups Chapter 22

Review Assessment-----

Chapter 3 SG 3.1 SG

3.2 SG 3.4 Chapter 3

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Chapter 3 Review

Physical and Chemical
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CHEMISTRY 705
SECTION 22-2 O
BJECTIVES Define
and relate the terms
radioactive decay and

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nuclear radiation

Describe the different types of radioactive decay and their effects on the nucleus Define the term half-life, and explain how it relates to the stability of a nucleus Define and relate the terms decay series, parent nuclide ...

~~Chapter 22 Review
Nuclear Chemistry~~

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Chapter 4 Answer Key

3 Practice Problems 13.

Use dimensional analysis and the given densities to make the following conversions. a.

14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/cm³.

$$(14.8 \text{ g}) \left(\frac{1 \text{ cm}^3}{2.34 \text{ g}} \right) = 6.32 \text{ cm}^3$$

b. 2.8 L of argon (Ar) to grams of argon.

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~~Chapter 4 Problem~~

~~Solving in Chemistry~~

~~Answer Key~~

1) No, it is not a proper

answer; you do not

know whether the

professor meant

homework problem

number 20 or 20

homework problems.

Twenty is a quantity

and does not make sense

without a unit. 3)

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Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

~~2.E: Homework~~

~~Chapter 2 Answers~~

~~Chemistry LibreTexts~~

Check the student submitted answers. 001 g in your data section.

00340 3 sig figs 14.

chapter 9 review

stoichiometry mixed

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