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ANS: The reversible reaction always attains equilibrium which proceeds both sides and never goes for completion. 2. The reaction CaCO3 [] CaO +CO2(g) goes to completion in lime because: a) Of the high temperature. b) CaO is more stable than CaCO 3. c) CaO is not dissociated.

Chemical Equilibrium Important Questions And Answers

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Chemical equilibrium is described as a dynamic process because there is a movement in which the forward and reverse reactions occur at the same rate to reach a point where the amounts or concentrations of the reactants and products are unchanging with time.

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Chemical Equilibrium Problems And Answer Key

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Answers To Chemical Equilibrium Study View Answer. Chemical equilibrium is established when the forward rate of reaction is equal to the reverse rate of reaction. a. TRUE b. FALSE. View Answer. For a reaction N₂ + O₂ to 2NO. If K_c... Chemical Equilibrium Questions and Answers | Study.com
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gives. Answers To Chemical Equilibrium Study Guide Answer and Explanation: A chemical reaction is a dynamic equilibrium. In chemical reactions when the system is at equilibrium, reactions are still occurring. However, the rate of forward reaction... A chemical equilibrium is a dynamic equilibrium. True ...

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Answer and Explanation: Chemical equilibrium is established d) when the rate of formation of products and the rate of formation of reactants becomes steady and equal, but does not stop the...

Chemical equilibrium is established: a) when ... - study.com

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14 D4 Which of the following describes all chemical equilibrium systems? A. The mass of the reactants equals the mass of the products. B. The species are present in the same ratio as in the balanced equation. C. The rate of the forward reaction equals the rate of the reverse reaction. D.

DYNAMIC EQUILIBRIUM STUDY GUIDE

Thus, the balanced chemical equation is expressed as 2BrN O → Br₂+2N O 2 B r N O → B r₂ + 2 N O. Now, we express the equilibrium constant based on the chemical equation and the corresponding...

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Balance the chemical reaction shown below and then write equilibrium expression. All reactants and all products are in the gas phase. N₂O₄ → 2NO₂ N₂O₄ → 2 N O₂

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Balance the chemical reaction shown below and then write equilibrium expression. All reactants and all products are in the gas phase. CO+F₂ →COF₂ C O + F₂ → C O F₂.

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Question: Experiment 10 Chemical Equilibrium: Determination Of An Equilibrium Constant INTRODUCTION In The Study Of Chemical Equilibria, Chemists Are Interested In Knowing Not Just Whether A Reaction Is Favored In The Forward Or In The Reverse Of Direction, But The Extent To Which It Is Favored. The Value Of The Equilibrium Constant, K, Provides This Information. ...

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