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15a: ANSWERS 1.

(a) Rate =  $k [H_2]$

$2[Br_2]^{0.5}$   $1.92 \times$

$10^{-3} M^{-1} s^{-1}$  (b)

2nd order,  $(2 + 0 =$

$2)$  2. (a) Rate =  $k$

$[O_2] [NO]^2$  (b)

$500 M^{-2} s^{-1}$  3. (a)

$(mol L^{-1})^{-2} s^{-1}$  (b)

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M-1 min-1 (c) g s-1

There are, of course, other ways to express these units correctly 4. Start with known concentrations of the two reactants and carry out the reaction.

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Mixing 1000 mL of  
an aqueous  
solution with 1000  
mL of water  
represents a A;  
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Complete the  
table. (5) Charge  
on most common  
ion Element Rb Cs  
At Se 2. Define,  
First Ionization

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Energy'. (2) 3.  
Using the metal  
magnesium as an  
example, write two  
separate equations  
to show the first

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. We are looking forward to teaching you during the 2019-2020 school year! Attached is the summer assignment. The first section is a summary of the information you should have

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learned in first-year chemistry. After the summary notes, you'll find several worksheets.

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