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~~TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry? 1. true 2.~~

~~true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon~~

~~dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles~~

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368 Chapter 11 □ Stoichiometry Section 11.1.1 Objectives Describe the types of relationships indicated by a balanced

chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting

substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

Chapter 11: Stoichiometry

alanle274. Chemistry Chapter 11 Stoichiometry. stoichiometry. mole ratio. limiting reactant. excess reactant. the study of

quantitative relationships between the amounts of.... in a balanced equation, the ratio between the number of moles.... a

reactant that is totally consumed during a chemical reaction....

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MANUAL Section 11.1 Defining Stoichiometry pages 368–372 Practice Problems pages 371–372 1. Interpret the following

balanced chemical equa-tions in terms of particles, moles, and mass. Show that the law of conservation of mass is

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CHAPTER STUDY GUIDE Date Class Stoichiometry Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

Human Resources Department – Dearborn Public Schools

In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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