

Chapter 12 Gaseous Chemical Equilibrium

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Chemical Equilibria • For a gaseous chemical equilibrium, more than one gas is present: • $aA(g) + bB(g) \rightleftharpoons cC(g) + dD(g)$ • To describe the state of this system, the partial pressures of all gases must be known • Using the Ideal Gas Law: • in a closed system, with fixed volume and temperature, the partial

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Chapter 12 Gaseous Chemical Equilibrium Outline 1. N₂O₄-NO₂ equilibrium system 2. The equilibrium constant expression 3. Determination of K 4. Applications of the equilibrium constant 5. Effect of changes in conditions on an equilibrium system Review of Liquid-Vapor Equilibrium • In Chapter 9 we examined the equilibrium that is

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Gaseous Chemical Equilibrium Chapter 12 2HI(g) H₂(g) + I₂(g) If pure HI(g) is placed in a sealed container at 52 C, H₂(g) and I₂(g) are formed. Then the reverse reaction can occur forming HI(g) The partial pressure of HI(g) drops, slowing the forward reaction, all the while, the partial pressures of H₂(g) and I₂(g) increase; increasing the rate ...

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the concentrations of the species remain constant over time. if the initial concentrations of the reactants and products change they will return to the same concentrations at equilibrium. only the equilibrium constant will remain constant. equilibrium partial pressure equation. $2A + 3B \rightleftharpoons C$. $K = \frac{(C_p)}{(A_p)^2 * (B_p)^3}$.

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Chapter 12: Gaseous Chemical Equilibrium Dynamic equilibrium- when the forward and backwards reactions of a system are taking place at the same rate. $(P_{\text{partial}})^V = (n)(0.08206)(T)$ [T is in Kelvin]. Dynamic equilibrium- when the forward and backwards reactions of a system are taking place at the same rate. $(P_{\text{partial}})^V = (n)(0.08206)(T)$ [T is

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7.1 EQUILIBRIUM IN PHYSICAL PROCESSES The characteristics of system at equilibrium are better understood if we examine some physical processes. The most familiar examples are phase transformation processes, e.g., solid \rightleftharpoons liquid liquid \rightleftharpoons gas solid \rightleftharpoons gas 7.1.1 Solid-Liquid Equilibrium Ice and water kept in a perfectly insulated

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