

## Chapter 4 Atomic Structure Answers

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### Chapter 4 Atomic Structure Answers

The nucleus of an atom consists of 8 protons and 6 neutrons. The total number of electrons present in an atom of this element is 8 What is the maximum number of electrons that can occupy the second energy level?

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The sum of the protons and neutrons in the nucleus of an atom.  $MASS\# = \# \text{ protons(atomic number)} + \# \text{ neutrons}$ .  $\#NEUTRONS = Mass \# - \#Protons(atomic number)$   $\#Protons = Mass \# - \#Neutrons$ . ISOTOPES: Are atoms of the same element that have different numbers of. neutrons and different mass numbers.

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### Chapter 4 Atomic Structure Answers

(i) Atomic number = number of protons + number of electrons (ii) Mass number = number of protons + number of neutrons (iii) Atomic mass = number of protons + number of neutrons (iv) Atomic number = number of protons = number of electrons (a) (i) and (ii) (b) (i) and (iii) (c) (ii) and (iii) (d) (ii) and (iv) Answer. Answer: (d) (ii) and (iv)

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