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quantum numbers
and electron

configuration ~~Chapter~~

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~~Atomic Structure and Periodic~~

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Properties of the Elements (Part 3)

~~Chapter 6 Electronic Structure Of~~

electron Light comes in packets of energy. Each packet runs into one electron. Each packet must have

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enough E to break electron loose. The rest of the energy goes into kinetic energy. Frequency tells us the E of each packet. I tells us how many packets/second we get. More packets, more current (more electrons knocked off). e^- $h\nu$ metal

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~~Structure of Atoms~~

Chapter 6 Electronic

Structure of Atoms

Chemistry, The

Central Science, 10th

edition Theodore L.

Brown; H. Eugene

LeMay, Jr.; and Bruce

E. Bursten Electronic

Structure of Atoms.

Waves Electronic

Structure of Atoms □

To understand the

electronic structure of

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atoms, one must understand the nature of

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Electronic Structure of Atoms. 4 a. False.

Electromagnetic radiation is capable of passing through

water. (That's why you can see through it.) b. True. c. True. d.

False. The glow from a fireplace and the energy within a microwave oven are forms of

electromagnetic

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radiation; the blast
from a foghorn is not.
6.

Atoms

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AP Chemistry

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electronic structure.

electromagnetic

radiation. wavelength.

frequency. the

arrangement of

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electrons in an atom or molecule. a form of energy that exhibits wavelike behavior as it travels. the distance between two peaks in a wave.

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structure ...~~

Chapter 6. Electronic
Structure and Periodic
Properties of

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Elements.

Introduction; 6.1

Electromagnetic

Energy; 6.2 The Bohr

Model; 6.3

Development of

Quantum Theory; 6.4

Electronic Structure of

Atoms (Electron

Configurations) 6.5

Periodic Variations in

Element Properties;

Chapter 7. Chemical

Bonding and

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Molecular Geometry.

Introduction; 7.1 Ionic
Bonding

~~6.4 Electronic~~

~~Structure of Atoms~~

~~(Electron~~

~~Configurations ...~~

6: Electronic Structure
of Atoms. In this
chapter, we describe
how electrons are
arranged in atoms
and how the spatial

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Arrangements of electrons are related to their energies. We also explain how knowing the arrangement of electrons in an atom enables chemists to predict and explain the chemistry of an element.

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An electron (mass = 9.109×10^{-31} kg) moving at 63.0 m/s *

Wavelengths of macroscopic particles are imperceptibly small and really have no physical

significance. 4.3×10^{-35} m * $(0.025 \text{ kg})(612 \text{ m/s})$ 6.63×10^{-34} kg m² / s 1.16×10^{-10} m $(9.109 \times 10^{-31} \text{ kg})(63.0 \text{ m/s})$

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6.6310kg m/s 5 31 34

2 36

Structure Of

Atoms

~~Chapter 6 Quantum~~

~~Theory and the~~

~~Electronic Structure of~~

~~Atoms~~

AP Chemistry

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- Chapter 6.

Electronic Structure of

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Review your notes

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from Honors or regular Chemistry for the sequence of atomic models and the evidence that allowed scientists to change the model. If you no longer have your notes, they are available

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at Harvard University.

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ELECTRONIC

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ATOMS 1 CONTENT

6.1 6.2 6.3 6.4 6.5 6.6

6.7 6.8 6.9 The wave

nature ...

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~~STRUCTURE OF~~

~~ATOMS 1 ...~~

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6 Electronic Structure
from CHM 71476 at
Arizona State
University. Electronic
Structure of Atoms
Chapter 6 How are
the electrons
distributed in atoms?
What are their relative

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6.1: The Wave Nature of Light.

Understanding the electronic structure of atoms requires an understanding of the properties of waves and electromagnetic radiation. A basic knowledge of the electronic structure of atoms requires an understanding of the properties of waves

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and electromagnetic radiation. A wave is a periodic oscillation by which energy is transmitted through space.

~~6.S: Electronic
Structure of Atoms
(Summary)~~

~~Chemistry ...~~

Electronic Structure.

This chapter is all about electronic

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structure—the arrangement and energy of electrons.. It may seem odd to start by talking about waves. However, extremely small particles have properties that can only be explained in this manner!

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Chapter 6 Electronic
Structure of Atoms.

Flashcard maker :

electronic structure:

The arrangement of
electrons of an atom
or molecule.

electromagnetic
radiation (radiant
energy) A form of
energy that has wave
characteristics and
that propagates

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through a vacuum at
the characteristic
speed of 3.00×10^8
m/s.

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Chemistry: The
Central Science (13th
Edition) answers to
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Exercises - Page 250

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6.30 including work
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Textbook Authors:

Brown, Theodore E.;

LeMay, H. Eugene;

Bursten, Bruce E.;

Murphy, Catherine;

Woodward, Patrick;

Stoltzfus, Matthew E.,

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Chapter 6 Lecture-

Electrons in Atoms 1.

Waves To

understand the

electronic structure of

atoms, one must

understand the nature

of electromagnetic

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radiation.

The distance between corresponding points on adjacent waves is the wavelength (λ).

2.

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Electrons in Atoms
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Electronic Structure □

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structure—the

arrangement and energy of electrons. □

It may seem odd to start by talking about waves. However, extremely small particles have properties that can only be explained in this manner!

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