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emiliosado176. Key Concepts: Terms in this set (79) Covalent bond. Atoms held together by sharing electrons. Molecule. A neutral group of atoms that are held together by covalent bonds (non-metals)

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CHAPTER 8 SOLUTIONS MANUAL Covalent Bonding  
Covalent Bonding Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240 – 247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH<sub>3</sub> H<sub>2</sub>S H<sub>2</sub>CO respectively, for single, double, and triple P — —  
2. H<sub>2</sub>S H<sub>2</sub>CO — H<sub>2</sub>S ...

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Chapter 8 Covalent Bonding and Molecular Structure 8-6 Example:  
PF<sub>3</sub> Example CO<sub>2</sub> Step 1: Count Valence Electrons Count the total number of valence electrons in the molecule or ion. Anions have extra electrons, so add 1 electron for each negative charge. Cations have a deficiency of electrons, so subtract 1 electron for each positive charge.

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chapter 8 test chemistry covalent bonding Flashcards and ... Section 8.2  
– The Nature of Covalent Bonding In ionic bonding, atoms transfer electrons to achieve noble gas configuration. In covalent bonding, atoms share electrons to achieve noble gas configuration.

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8.8: Strengths of Covalent Bonds Bond enthalpy is the enthalpy change for the breaking of a particular bond in one mole of a gaseous substance o The designation  $D(\text{bond type})$  is use to represent bond enthalpies Many important bonds such as C—H bond only exist in polyatomic molecule

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In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). 12. In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). But rather than losing or gaining electrons, atoms now share an electron pair. 13. In covalent bonding, atoms still want to achieve a noble gas ...

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