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Ionization Energy,

Atomic Radius—

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How to calculate
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Levels, Energy

Sublevels, Orbitals,

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Exclusion Principle

Drawing \u0026

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*Trends in the periodic
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**Quantum Numbers -
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8-1 Figure 8.6 Orbital Filling Order . Chapter 8 Electron Configuration And General

& Chemical Periodicity

. One additional quantum number is needed to describe a property of an electron in an atomic orbital. 4. Spin

Quantum Number, m_s . $m_s = +\frac{1}{2}$ or $-\frac{1}{2}$

2. An electron in an

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orbital is described by its set of FOUR quantum numbers. “shell”: orbitals with the same

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PERIODICITY | ?????

?????????? -

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END-OF-CHAPTER

PROBLEMS 8.1

Elements are listed in the periodic table in an ordered, systematic way that correlates with a periodicity of their chemical and physical properties.

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CONFIGURATION

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Quiz 3. 1. In

predicting the electron

configuration of the

elements by the

Aufbau Principle, to

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which sublevel is one adding electrons in traversing from element 39, Y, to element 48, Cd? a) 4f b) 5p c) 4d d) 5d. 2.

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tanzaniabj. Terms in

this set (8) Electron

configuration. is a

particular distribution

of electrons among

the available

subshells. subshells.

consists of a group of

orbitals having the

same n (energy level)

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and l (s,p,d,f)

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8-30 Sample Problem

8.2 Determining

Electron

Configurations

PROBLEM: Using the

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periodic table on the inside cover of the text (not Figure 8.10 or Table 8.3), give the full and condensed electron configurations, partial orbital diagrams showing valence electrons only, and number of inner electrons for the following elements:

(a) potassium

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CONFIGURATION AND THE PERIODIC TABLE. The electrons in an atom fill from the lowest to the highest orbitals. The knowledge of the location of the orbitals on the periodic table can greatly help the writing of electron configurations for large atoms. 15.

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electron spin. 8-2

Summary of Quantum
Numbers of

And General

Chemistry

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Chemical Periodicity.
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magnetic quantum

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number, m_s .

describes the angular momentum, or spin, of an electron. values of m_s $+1/2$ and $-1/2$.

Pauli exclusion principle.

Chapter 8: Electron Configuration and Chemical Periodicity

...

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Chapter 8: Periodic
Trends Atoms gain,
lose, or share

electrons to obtain a
noble?gas electron
configuration Noble

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Gas Electron

Configuration =
 ns^2np^6 = an octet, 8
electrons in the
valence shell (highest
n value) This is a very
stable arrangement
because it ...

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Configuration and
Chemical Periodicity

8.1 Elements are listed in the periodic table in an ordered, systematic way that correlates with a periodicity of their chemical and physical properties. The theoretical basis for the table in terms of atomic number and electron configuration

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does not allow for an "unknown element" between Sn and Sb.

8.2 Today, the elements are listed in order of increasing atomic number.

Chapter 8 Electron Configuration and Chemical Periodicity

...

where [He] represents

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the two-electron core that is equivalent to He's electron configuration. The square brackets represent the electron configuration of a noble gas. This is not much of an abbreviation.

However, consider the abbreviated electron configuration for W, which has 74

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electrons: W: [Xe]6s 2
4f 14 5d 4

Configuration

And General

Organization of

Electrons in Atoms –
Introductory ...

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Educators. ML RP + 4
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Chapter Questions.

02:44. Problem 1

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What would be your reaction to a claim that a new element had been discovered and it fit between tin (Sn) and antimony (Sb) in the periodic table. ML Marcus L. ...

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ATOMIC ELECTRON
CONFIGURATIONS
AND PERIODICITY 1

Chapter 8 ATOMIC
ELECTRON

CONFIGURATIONS
AND PERIODICITY 2

Arrangement of
Electrons in Atoms.

Electrons in atoms
are arranged as ;

SHELLS (n)

SUBSHELLS (l)

ORBITALS (ml) 3

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Arrangement of
Electrons in Atoms.
Each orbital can be
assigned no more
than 2 electrons!

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011f7c7de