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~~Heck is an Acid or
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What Are Salts? |
Acids, Bases

Alkali's |
Chemistry |

FuseSchool Ionic
Product of Water
(K_w) : Self

Ionization of water

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~~Chapter 8 Solutions~~
~~Acids Bases~~
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solution. 8.3

Properties of Acids and Bases Some general properties of acids include sour taste, reactivity with metals, and ability to produce color changes in indicators. □ An acid is a compound

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Chapter 8

Solutions, Acids
Bases, Essment

that produces hydronium ions (H_3O^+) when dissolved in water.

□ An indicator is any substance that changes color in the presence of an acid or base.

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substance that is
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solvent to make a
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substance in which
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dissociation.

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substance that is

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which a solute is

dissolved to form a

solution.

~~Chapter~~

~~8 Solutions, Acids,~~

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~~& Bases, Chapter~~

~~8: Solutions ...~~

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An acid produces hydronium ions in solution.

Acceptable answers include hydrochloric acid, citric acid, and acetic acid. A base produces hydroxide ions in solution. Sodium hydroxide.

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Bases, Acids
Formation of

Solutions (pages
228–234) This
section explains
the parts of a
solution, the
processes that
occur when

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Solutions, Acids,
and Bases Section

8.4 Strength of

Acids and Bases

(pages 246–249)

This section

explains how to

describe acids and

bases in terms of

both concentration

and strength.

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Solutions, Acids

Bases Assessment
(page 246)

Comparing and

Contrasting As you
read, complete the
diagram by

comparing and
contrasting acids
and bases. For
more information

on

~~Chapter 8~~

~~Solutions, Acids,~~

Page 24/44

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Chapter 8

~~and Bases Section~~

~~8.4 Strength ...~~

Chapter 8.

Solutions, Acids &
Bases. Dissolving:
Dissociation,
Dispersion or
Ionization

Substances can
dissolve in water
by three ways: 1.
dissociation 2.
dispersion 3.
ionization

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Dissociation is the process by which an ionic compound separates into ions

Ex: NaCl Dispersion

is the process by which a compound breaks into small pieces that spread throughout the

water Ex: sugar dissolves in water

Ionization is the process by which

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neutral molecules
gain or lose
electrons (forming
ions) Ex ...

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Hall Physical
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in Action~~

Chapter 8 Acids,
Bases, and Acid-
Base Reactions An
Introduction to
Chemistry by Mark

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Bishop. Chapter
Map. ... solution is
neutral in the
acid/base sense.

pH □To avoid the
small numbers
associated with
describing acidic
and basic solutions
in terms of mol/L,
pH is defined as

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Bases, and Acid~~

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~~Base Reactions~~
Solutions Acids

Bases Essment
This section begins with a review of

acids, followed by the following for bases: (1) it states the Arrhenius definition of base, (2) it provides you with the information necessary to identify strong and weak bases, and

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(3) it describes the changes that take place when one weak base (ammonia) dissolves in water (Figure 8.2).

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Solutions Acids and Bases July 30,

2020 1. Definition of an Arrhenius

(1887) acid and

base: 2. Definition of a Bronsted-

Lowry (1923) acid and base: 3. Write

the neutralization reaction for: HNO_3

(aq) + NaOH (aq) 4.

Write the reaction of vinegar (acetic

acid) and baking

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Soda (sodium bicarbonate): 5.

Write the reaction of limestone (CaCO_3) and HCl . 6.

~~20200730 Chapter 8a Worksheet Acids and Bases.doc ...~~

Chapter 8
vocabulary
(Solutions, Acids,
and Bases) 8.1-
Formation of

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Solutions Ionization

The process in which neutral molecules gain or lose electrons.

Solute A substance whose particles are dissolved in a solution. Kool-Aid

Dispersion The breaking into the small pieces that

~~Chapter 8~~

Page 34/44

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~~vocabulary~~
~~(Solutions, Acids,~~
~~Bases, Essment~~
~~and Bases) by ...~~

Base: A compound that produces hydroxide ions when dissolve in water; a proton acceptor:

Neutrilation: A chemical reaction between and acid and a base: Salt: An ionic compound

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Solutions, Acids
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formed when an acid reacts with a base: pH: A measure of the hydronium ion concentration of a solution: Buffer: A solution that is resistant to large changes in pH ...

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Acids, and Bases

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chapter of this

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helps students

learn the essential

lessons associated

with acids, bases

and solutions. Each

of these...

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Edit. ... Which
methods will
distinguish

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between equimolar solutions of a strong base and a strong acid? I. Add magnesium to each solution and look for the formation of gas bubbles II. Add aqueous sodium hydroxide ...

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Chapter 8, Bases Assessment

Solutions, Acids,
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Particles that are dissolved in a solution. The solution in which a solute will dissolve in. The process in which an ionic compound separates into ions as it dissolves.

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Solutions, Acids, Bases, Essential

Breaking down into small pieces that spread out through the water.

~~Quia Chapter 8,
Solutions, Acids,
and Bases~~

In Chapter 8.6, we defined acids as substances that dissolve in water to produce H^+ ions, whereas bases

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~~Solutions Acids~~
~~Bases Essment~~
were defined as substances that dissolve in water to produce OH^- ions. In fact, this is only one possible set of definitions.

~~Chapter 8.07: Acid~~

~~Base Reactions~~

~~Chemistry~~

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Solution: The

reaction of the acid

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+ base gives a product of salt + water, which is considered as neutralization reaction.

Examples: $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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