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Chemical Bonding SECTION 4 SHORT
ANSWER Answer the following questions
in the space provided. 1. b In metals, the
valence electrons are considered to be (a)
attached to particular positive ions. (c)
immobile. (b) shared by all surrounding

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atoms. (d) involved in covalent bonds. 2. a
The fact that metals are malleable and
ionic crystals are brittle is best

6 Chemical Bonding

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Worksheets - Kiddy Math

Draw diagrams (outer electrons only) to show the bonding in the following covalent molecules. Draw a before and

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after bond picture for each on another sheet of paper (your drawings should be similar to the example above):: Hydrogen fluoride \square HF. Chlorine \square Cl₂. Oxygen \square O₂. Nitrogen \square N₂. Silicon tetrachloride \square SiCl₄. Ammonia \square NH₃

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Worksheet Chemical Bonding Ionic and Covalent Answers Part 2 Along with Lovely Ionic Bonding Worksheet Answers Best Chemical Bonds. The fact that a molecule will react when it has an ionic bond is not new to us. It is something that all chemical scientists know already, and

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just as many professionals in other fields
that deal with different ...

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Chemistry: Chemical Bonding Activity.
Introduction. When atoms bond together

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to form ionic compounds, they will not combine with just any other atom. For example, two atoms that will never form an ionic bond are a sodium atom (Na) and a potassium atom (K). This is because both Na^{1+} and K^{1+} are cations, or positively-charged ions.

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Chemical Bonding Activity

WORKSHEET: Ionic vs. Covalent! Ionic Bond between a Metal and Non-Metal (M + NM) Covalent Bond between a Non-Metal and Non-Metal (NM + NM)

Determine if the elements in the following compounds are metals or non-metals.

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Describe the type of bonding that occurs in the compound.

WORKSHEET: Ionic vs. Covalent!

REMEMBE Ionic Bond Covalent ...

water covalent bond □ REFERENCE

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from the attraction between the valence
electrons and of different atoms.

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bonds that are made by the two elements

...

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1. The valency of an element is _____ (a)
the combining capacity of one atom of it
(b) the number of bonds formed by its one
atom (c) the number of hydrogen atoms

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that combine with one atom of it (d) all the
above Answer. (d)

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