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engineering (CRE): Chemical reaction engineering is that engineering activity concerned with the exploitation of chemical reactions on a commercial scale. Its goal is the successful

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Theory. The objectives of the development that follows are to give the reader insight as to why the rate laws depend on the concentration of the reacting species (i.e., $r_A = k C_A C_B$) and why the temperature dependence is the form of the Arrhenius law,

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$k = A e^{-E_a/RT}$. To achieve this goal we consider the reaction of two molecules in the gas phase

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look at the formation of ammonia (NH₃) from nitrogen and hydrogen (see section Chemical equilibrium of the thermodynamics chapter). This reaction follows the equation: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ (1) $H_0 = 92 \text{ kJ mol}^{-1}$ $S_0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$

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state, that is, 1 atm for gases and
thus $K = \frac{P_{\text{N}_2}^{1/2} P_{\text{H}_2}^{3/2}}{P_{\text{NH}_3}}$ Use
of the Lewis and Randall rule
gives:

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which of the following cases does the reaction go farthest to completion? a. $k = 10^2$ c. $k = 10$
b. $k = 10^{-2}$ d. $k = 1$ 2. If "n" is the order of reaction, then unit of rate constant is a. $1 / (\text{time})$
(concentration)ⁿ⁻¹ b. (time)⁻¹
(concentration)ⁿ⁻¹ c.

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