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Sample Problem #2. Write the chemical equation for the following reaction:

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Hydrochloric acid and solid sodium hydrogen carbonate react to produce aqueous sodium chloride, water, and carbon dioxide. Hint: Acids will always be aqueous unless otherwise stated.  $\text{HCl (aq)} + \text{NaHCO}_3\text{(s)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)} + \text{CO}_2\text{(g)}$

## ~~Chapter 11: Chemical Reactions~~

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GUIDED PRACTICE PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction. Include appropriate symbols from table 11.1. Analyze Step 1. Identify the

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Chemical Reactions Guided Practice Problem CHEMICAL REACTIONS (pages 330-339). This section GUIDED PRACTICE PROBLEM 2 (page 324). 2. possibly a change of state. Essential Questions and Answers. 1. What are the types for energy? Chemical Energy is energy caused by chemical reactions. Quiz: Chemical Reactions. 1/8. During chemical reactions, atoms.

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13) Using the equation from problem #12, determine the mass of aluminum acetate that can be made if I do this reaction with 125 grams of acetic acid and 275 grams of aluminum hydroxide. 14) What is the limiting reagent in problem #13? 15) How much of the excess reagent will be left over in problem #13 after the reaction is complete?

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Name: \_\_\_\_\_ Unit 11 Honors Chemistry Notes Thermodynamics – the study of \_\_\_\_\_. Thermochemistry is the study of how heat changes during phase changes and chemical reactions. Types of Energy + Quantifying Energy (ability to do work or produce heat) Calories Joules Law of Conservation of Energy – 1<sup>st</sup> Law of Thermodynamics ex) Potential Energy stored in gasoline bonds Kinetic energy in ...

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stoichiometry grams-grams problems: 1) Using the following equation:  $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$  How many grams of sodium sulfate will be formed if you start with 200.0 grams of sodium hydroxide and you have an excess of sulfuric acid?  
2) Using the following equation:

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