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Chem. 112 Final Exam Practice 1 Part 4 1) The decomposition of hydrogen iodide follows the equation  $2\text{HI}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \text{I}_2(\text{g})$  and it is second order with a rate constant of  $1.6 \times 10^{-3} \text{ L/mol sec}$  at  $700^\circ \text{C}$ . If the initial concentration of HI in the container is  $3.4 \times 10^{-2} \text{ M}$ , how many seconds will it take for

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*CHAPTER 12 REVIEW Solutions*

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CHAPTER 13 REVIEW Solutions SECTION 13-2 SHORT ANSWER Answer the following questions in the space provided. 1. Following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the liquid. b.

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