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GRAMS Molar mass (grams per mole) # MOLES 6.02 X 10²³ atoms or molecules per mole # MOLECULES 22.4 liters of gas per mole # LITERS (at STP) EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP? # liters Ne = 40.36 g Ne x 1 mole Ne x 22.4 L Ne = 44.80 L Ne 20.18 g Ne 1 mole Ne EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP? # molecules CO₂ = 56.0 L CO₂ x 1 mole CO₂ x 6.02 X 10²³ molecules CO₂ = 1.51 x 10²⁴ molecules CO₂ 22.4 L CO₂ 1 ...

CHEMISTRY B- MOLES PACKET NAME: HR: _____ PAGE 1

CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 3 CALCULATE THE MOLAR MASS FOR THE FOLLOWING COMPOUNDS OR DIATOMIC ELEMENTS. SET UP EACH PROBLEM AS SHOWN IN THE EXAMPLE ABOVE. INCLUDE UNITS (G/MOL) 1. water H₂O 2. calcium chloride CaCl₂ 3. copper(II) sulfate CuSO₄ 4. silver nitrate AgNO₃ 5. sulfuric acid H

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 6CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBERNow that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you cancombine these two definitions into one problem.EXAMPLE: How many molecules are there in 90.1 grams of water?2 H = 2 x(1.01) = 2.02O = 1 x (16.00) = 16.00H 2 O= 18.02 g? molecules H 2 O = 90.1 g H 2 O x 1 mole H 2 O x 6.02 x 10 23 molecules H 2 O = 3.01 x 10 24 molecules H 2 ...

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