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empirical formula 2)

calculate molar mass of  
empirical formula 3)

divide the molar mass of  
the molecular formula

(given) by the molar

mass of the empirical

formula 4) multiply all

subscripts in the

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empirical formula by the  
rounded

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What is a chemical

change? 4. What is the

periodic table? 5.

Complete the table

below.

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relationships between

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used and products

formed by a chemical

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Calculate the final pressure of a sample of carbon dioxide that expands reversibly and adiabatically from 57.4 kPa and 1.0 dm<sup>3</sup> to a final volume of 2.0 dm<sup>3</sup>. Take  $\gamma = 1.4$ .

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are many benefits to  
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mind that you are trying  
to ... Round your final  
answer to the

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appropriate place. 2. An object has a mass of 19.355g and a volume of 11.5 mL. a. The measurement with the least number of ...

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Electronegativities of bonding atoms are the same (as they are in diatomic molecules) or close to the same, they SHARE electrons.

Bonds formed when atoms share electrons are called Covalent Bonds .

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the brain to think  
augmented and faster  
can be undergone by  
some ways.

Experiencing, listening

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to the other experience,  
adventuring, studying,  
training, and more  
practical comings and  
goings may incite you to  
improve.

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