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Recording] How To
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Review Choosing
Between
SN1/SN2/E1/E2
Mechanisms 01 -
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SEMESTER 2 FINAL
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DEBATE

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Stoichiometry
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and more with
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tryGas
LawsSolutionslons
in Aqueous
SolutionsAcids and
BasesOxidation

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and Reduction
Reactions (and Elec
trochemistry) Practi
ce Problems (Skills)
Terms in this set
(59) Mole. The
amount of a
substance whose
mass is equal to
the atomic mass in
grams. Avogadro's
number.

Chemistry

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Honors Chemistry

Name Semester 2

Final Exam Period

PART □ □ □ A □.

MULTIPLE CHOICE

Date (PRACTICE)

You may use a pencil, eraser, and scientific calculator to complete the test. You will be

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Final Exam
Study Guide
Answers

given a periodic table. Useful formulas and the solubility rules are listed...

Chemistry
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Exam Multiple
Choice
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Exam Review
DRAFT. 11th grade.

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67 times.

Chemistry. 74%
Final Exam
average accuracy.
Study Guide
a year ago.

Answers. 1. Save.

Edit. Edit.

Chemistry

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Exam Review

DRAFT. ... If the
saturated solution
of $K_2Cr_2O_7$ was
cooled from 80 to
60 °C, how much

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Solute would fall
out? answer
choices . 20 g. 40
g. 15 g . 30 g. Tags
Answers

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Fall 2020 Study
Guide I will print

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Copies of this and stick in a labeled box in the circle in front of the school near D hallway for any student to pick up on 12/3 – 12/4. Students in person can pick up in class on same days.

Final Exam Format
(subject to change)-1 hour: 3 FRQs, due on time.

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Students in person will just write on the test itself.-30 minutes - 40 ...

Answers

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Learning goals
Checklist As you

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prepare for your
final exam use the
following list to
monitor your
progress. Only
check off a learning
goal when you are
confident of your
ability to complete
the task described.
Unit 0: Being a
Chemist Recognize
and apply safe
practices when

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working in the lab.

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Exam in this class
is the Michigan
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Grades will be

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based on achieving the Department of Education recommended passing rate of 70% (For Example - correctly answering 70% of the questions = 100% recorded in Powerschool as your Exam score)

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easy to get the
grade you want!
minimum amount
of energy needed
by the colliding
molecules in a
chemical reaction
to break the bonds
of the reactants.

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er-Final-Exam-
Review-Sheet 2/3

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Physical Chemistry
II 3 Semester

Credits

(4120/6120)

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(Semester 2) Part
A: Multiple Choice
and Numerical
Response 1. Which
of the following
equations is

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associated with the largest energy change per mole of fluorine? A. $F_2(g) \rightarrow F_2(l)$ B. $19\ 9\ F + 4\ 2\ He \rightarrow 1\ 0\ n + 22\ 11\ Na$ C. $2\ F_2(g) + 2\ H_2O(l) \rightarrow O_2(g) + 4\ HF(aq)$

Final Exam Practice
Test (Semester 2)
CHEMISTRY I
HONORS – FINAL

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EXAM REVIEW

ANSWER KEY 1. $1s$
 $2s2p63s23p64s2$
 $3d104p3$ 2. bottom-
left 3. chemical 4.
chemical 5.
physical 6. low PE
= high stability 7.
MgF 2 8. N 2 O 5 9.
Na 2 SO 4 10. H 3
PO 4 11. potassium
nitrate 12.
hydrobromic acid
13. sulfur trioxide

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14. iron(III) chloride
15. polar (see diagram)
16. nonpolar (see diagram)
17.

CHEMISTRY I
HONORS FINAL
EXAM REVIEW
6.1.3 Final Exam:
Semester 2 Exam
Chemistry Sem 2
Points Possible:
100 Name: Date:

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3/4/2020 In this test, you will apply what you have learned in the semester. Answer the questions below. Submit this to your teacher when you are done. Show all your work on this test.

6.1.3 Final Exam_
Semester 2

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Exam.docx - 6.1.3

Final Exam ...

Chemistry

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Review. Terms in

this set (42)

Explain how a molecular formula and an empirical formula are related? Molecular formula is a multiple (n) of the EMPIRICAL formula.

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What is the unit of measurement for molar mass?
Grams.

Answers

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□□question Mole
answer The amount

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Answers

of a substance whose mass is equal to the atomic mass in grams question Avogadro's number.

answer 6.022×10^{23}

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of most gases is

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expansion, fluidity,
compressibility,
and diffusion

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o Know the
difference between
directly and
inversely

proportional. o

Know that the
pressure and

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volume of a gas
are inversely
proportional. o
Know how to use
Boyle's Law to
solve for the initial
or final pressure or
volume of a.

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Answers (Seniors
) REMINDER:

DON'T FORGET

YOUR NON-

GRAPHING

SCIENTIFIC

CALCULATOR THE

DAY OF THE EXAM!

You will be

provided with a

periodic table with

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the polyatomic ion chart and equations that appear on the last page of this packet.

Your organization will be assessed prior to the Final Exam, so be sure to have all of your Semester 1 and Semester 2 materials with you once review

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