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1st Law of Thermodynamics

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Basic concepts and definitions - Part 1

The Energy Equation | First Law of Thermodynamics | Work | Heat Transfer | Internal Energy | Enthalpy

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about 1 kg/m³ so we get m = ρV = 1 kg/m³ × 0.001 m³ = 0.001 kg A more accurate value from. Table A.5 is ρ = 1.17 kg/m³ at 100 kPa, 25°C.

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The equation for a liquid is: v = Constant = v_o. If you include that v increases a little with T then: v = v_o + C (T - T_o) where C is a small constant with units m³/kg-K. Full file at <http://TestBankSolutionManual.eu/Solution-for-Fundamentals-of-Thermodynamics-8th-Edition-by-Borgnakke-International-version->

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Solution: m = ρV = 110 kg/m³ × 100 000 m³ = 11 × 10⁶ kg Just to put this in perspective a power plant that generates 2000 MW by burning coal would make about 20 million tons of carbon-dioxide a year. That is 2000 times the above mass so it is nearly impossible to store all the carbon-dioxide being produced.

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Solution: $Hg : L = 1 \text{ m}; \rho = 13\,580 \text{ kg/m}^3$ from Table A.4 (or read Fig 2.7) The pressure difference ΔP balances the column of height L so from Eq.2.2 $\Delta P = \rho g L = 13\,580 \text{ kg/m}^3 \times 9.80665 \text{ m/s}^2 \times 1.0 \text{ m} \times 10^{-3} \text{ kPa/Pa} = 133.2 \text{ kPa}$ Sonntag, Borgnakke and van Wylen 2.14 You dive 5 m down in the ocean.

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