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~~Answers~~ *How to Pass Chemistry Stoichiometry Problems I (Mixed Stoich Review)*

Balancing Chemical Equations

Practice Problems **Solution**

Stoichiometry - Finding Molarity,

Mass \u0026amp; Volume Mole Ratio

Practice Problems 1 and 2 Step

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Limiting Reactant and Excess Reactant

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Stoichiometry Made Easy: The Magic

Number Method Stoichiometry Made

Easy: Stoichiometry Tutorial Part 1

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~~Answers~~ Mass conversion problem

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How to Find Limiting Reactants | How to Pass Chemistry Solving Solution Stoichiometry Problems Limiting Reactant Practice Problem

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Reactant Stoichiometry & Moles*

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Using Molarity as a Conversion Factor
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CHAPTER 9 REVIEW Stoichiometry

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MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of

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Chapter 9 Stoichiometry Review
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Modern Chemistry 2 Stoichiometry
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MIXED REVIEW SHORT ANSWER

Answer the following questions in the

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Answer provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ _____ a. What is the value of the coefficient x in this equation?

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CHAPTER 9 REVIEW Stoichiometry

MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation?

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40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O 2:1 mol H $2O$ c. What is the mole ratio ...

Chapter 9 Stoichiometry Review

Answers Section 2

ANSWER KEY. Mixed Stoichiometry

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Answers

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. a). How many moles of H_2 would be required to produce 5.0 moles of water? 5.0 moles water. b). What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ? 432g H_2 .

2. $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$. a). Balance this equation. Look above.

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b). Answers

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SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is Page 11/18

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ANSWERS SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is Page 11/18 ...

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Worksheet Answers Mixed

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Stoichiometry: Mixed Problems (KEY)

1) $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ What volume of NH_3 at STP is produced if 25.0 of N_2

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Answers
is reacted with an excess of H₂? 3 3 3
2 3 2 2 2 40.0L NH₃ 1mol NH₃ 22.4L NH₃
1mol N 2mol NH₃ 28.0g N 25.0g N
1mol N x x x = 2) 2KClO₃ ? 2KCl +

Mixed Stoichiometry Worksheet
Answers

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Stoichiometry Review Answers 1. a.
Na₃PO₄ b. Ca(NO₃)₂ Na = 3 mol x
22.99 g/mol = 68.97 g Ca = 1 mol x
40.08 g/mol = 40.08 g P = 1 mol x
30.97 g/mol = 30.97 g N = 2 mol x
14.01 g/mol = 28.02 g O = 4 mol x
16.00 g/mol = 64.00 g O = 6 mol x
16.00 g/mol = 96.00 g 163.94 g 164.10

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g c. $\text{Ca}_3(\text{PO}_4)_2$ d.

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Answer the following questions in the

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space provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation?
40.07 g/mol b. What is the molar mass of C_3H_8 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio of O_2 to H_2O

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Mixed Stoichiometry Problems. 1.
Hydrogen and oxygen react under a
specific set of conditions to produce
water according to the following: 2H_2
(g) + O_2 (g) \rightarrow $2\text{H}_2\text{O}$ (g) A. How many

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Answers
moles of hydrogen would be required in order to produce 5.0 moles of water? B. How many moles of oxygen are required to produce 436 L of water vapor?

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Answer Key

Stoichiometry. Stoichiometry is the practice of predicting the amount of

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product or reactant in a chemical equation based on a known amount of one of the other products or reactants. The ability to do these calculations is the culmination of all of the basic skills learned throughout the first semester.

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Stoichiometry SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ _____ a. What is the value of

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the coefficient x in this equation?

40.07 g/mol h. What is the molar mass of C₃H₄? 2 mol O₂:1 mol H₂O c.

Date. FCHAPJ REV[EW.

Stoichiometry Review Worksheet.

1) Using the following balanced

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Answers: $2 \text{NaOH (aq)} + \text{H}_2 \text{SO}_4 \text{ (aq)} \rightarrow 2 \text{H}_2 \text{O (l)} + \text{Na}_2 \text{SO}_4 \text{ (aq)}$ How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid? 2) Using the following balance equation: $\text{Pb(NO}_3)_2 \text{ (aq)} + 2 \text{Li}_2 \text{SO}_4 \text{ (aq)} \rightarrow \text{Pb(SO}_4)_2$

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(s) + 4 LiNO₃ (aq)

Stoichiometry Practice Worksheet

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W/ answers Website Upload 14. 2 Na

+ 2 H₂O ? 2 NaOH + H₂ 15. 2 Al₂O₃ +

3 HgCl₂ ? 2 AlCl₃ + 3 HgI₂ 16. 3

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Answers
Ca(OH)₂ + 2 H₃PO₄ ? Ca₃(PO₄)₂ + 6 H₂O 17. 3 AgNO₃ + K₃PO₄ ? Ag₃PO₄ + 3 KNO₃ 18. C₃H₈ + 5 O₂ ? 3 CO₂ + 4 H₂O Use the law of conservation of mass to determine the missing reactant in the equation given ...

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Unit 6: Reactions and Stoichiometry
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the following questions in the space

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Answers provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ _____ a. What is the value of the coefficient x in this equation? _____ b.

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