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CHAPTER 9 REVIEW Stoichiometry SECTION 3

PROBLEMS Write the answer on the line to the left.

Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Describe the relationship between Avogadro's number and a mole of a substance. A mole is defined as the amount of a substance in 6.022×10^{23} (Avogadro's number) particles.

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Review page 301 1- The branch of chemistry that deals

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with mass relationships in compounds and in chemical reactions. $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$ 2- a) $2 \text{ mol HgO} \rightarrow 2 \text{ mol Hg} + 1 \text{ mol O}_2$...

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6 Chemical Bonding

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