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## Mole Lab Answers

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Chemistry Experiment: The Mole — What Does  
it Look Like? (Part 1) Molarity Practice  
Problems Target Mole Lab Empirical Formula of  
Magnesium Oxide Post-Lab

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*Stoichiometry Using Molarity as a Conversion*

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*Tutorial: Step by Step Video + review*

*problems explained | Crash Chemistry Academy*

**Finding and Calculating an Empirical Formula**

**of a Compound | How to Pass Chemistry** ~~Using~~

~~Avogadro's Number | How to Pass Chemistry~~

Molarity Made Easy: How to Calculate Molarity

and Make Solutions The Mole **Interconverting**

**Masses, Moles and Numbers of Particles -**

**Chemistry Tutorial Converting Between Moles,**

**Atoms, and Molecules** ~~What is an equivalent? |~~

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~~Lab values and concentrations | Health \u0026amp; Medicine | Khan Academy~~ Molar Volume of H<sub>2</sub>

Lab An Open Window | Critical Role | Campaign 2, Episode 114 **Mole Conversions Made Easy:**

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ll iht eered Write the actual mass here 0.98

x (correct answer) 0.95 x (correct answer)

0.90 x (correct answer) 1.02 x (correct

answer) 1.05 x (correct answer) 1.10 x

(correct answer) Correct answer: (actual mass

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divided by 63.546 times 6.022e23) 4 3 2  
Figure 1. Procedure 1.

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ll iht eered Write the actual mass here 0.98  
x (correct answer) 0.95 x (correct answer)  
0.90 x (correct answer) 1.02 x (correct  
answer) 1.05 x (correct answer) 1.10 x  
(correct answer) Correct answer: (actual mass  
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Copper (II) Sulfate Pent hydrate,

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Solutions—Alum Moles Lab Activity 5:  
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The fastest way to obtain a relative mass of beans would be to count the beans. The fastest way to obtain a mole of beans would be to weigh them. (At least in principle. The mass of a mole of beans would be incredibly large- on the order of  $10^{22}$  g.) Part III. All atomic masses agree with the relative



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masses to three significant figures.

*Laboratory Activity 1: Teacher Notes  
Continued*

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Have a conceptual understanding of the mole

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as the method of “counting” items and finding the mass of items that can't be seen. 2. Be able to calculate the number of items (molecules, atoms, ions, and formula units) if given the number of moles. 3. Be able to calculate the number of moles if given the number of items.

*#20 Introduction to the Mole - Terrific Science*

Answer # target gas law lab continue 5 216  
linn cientiic inc all ights esered target gas  
law lab worksheet many chemical reactions  
involve gaseous reactants or products the

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volumes of gases that react or are produced depend entirely on the moles of the gas and of course on the pressure and

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