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(p.429 in the Textbook)

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Solution: OH^- OH^- OH^- 0.50 mmol mL – – – = x = $n_c V$ x 25.0 mL
 OH^- n – = 12.5 mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = = n
 V 0.10 mmol HCl(aq) /mL V = 120 mL 2. (a) Given: V HCl(aq) = 25.00 mL; [HCl(aq)] = 0.350 mol/L; V NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H⁺(aq)], pH Analysis: Before

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titration, amount of H⁺ is: H(aq) H(aq) H(aq) 0.350 mmol mL ++
+= x = nc V

Section 8.7: Acid – Base Titration Tutorial 1 Practice, page 547

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Solution: $Q = [\text{H}_2(\text{g})][\text{I}_2(\text{g})][\text{HI}(\text{g})]^2 = (0.040)(0.010)(0.14)^2$
 $Q = 0.020$ Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given: $[\text{HI}(\text{g})] = 0.20 \text{ mol/L}$; $[\text{H}_2(\text{g})] = 0.15 \text{ mol/L}$; $[\text{I}_2(\text{g})] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ; direction of reaction to reach equilibrium Solution: $Q = [\text{H}_2(\text{g})][\text{I}_2(\text{g})][\text{HI}(\text{g})]^2 = (0.15)(0.090)(0.20)^2$ $Q = 0.34$

Section 7.5: Quantitative Changes in Equilibrium Systems ...

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