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Write balance equations for
the following redox

reactions: a. $\text{NaBr} + \text{Cl}_2$

$\text{NaCl} + \text{Br}_2$ b. $\text{Fe}_2\text{O}_3 + \text{CO}$

$\text{Fe} + \text{CO}_2$ in acidic solution

c. $\text{CO} + \text{I}_2\text{O}_5 \rightarrow \text{CO}_2 + \text{I}_2$

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in basic solution Hint;
Write balanced equations for
the following reactions:

Hint. a. $\text{Cr(OH)}_3 + \text{Br}_2 \rightarrow \text{CrO}_4^{2-} + \text{Br}^-$ in basic solution.

b. $\text{H}_2\text{O}_2 + \text{SbO}_2 \rightarrow \text{SbO}_3^{2-}$ in basic solution Hint

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Practice Problems: Redox Reactions

Return to Redox menu.

Problem #1: $\text{Cr}^{2+} + \text{Fe}^{3+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{2+}$.

Solution: 1) Balanced half-reactions:
 $6e^- + 14\text{H}^+ + \text{Cr}^{2+} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2$

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0. $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$. 2)
Equalize the electrons: $6e^-$
 $+ 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow$
 $2\text{Cr}^{3+} + 7\text{H}_2\text{O}$. $6\text{Fe}^{2+} \rightarrow$
 $6\text{Fe}^{3+} + 6e^-$ <--- multiplied
by a factor of 6.

Balancing redox reactions in

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acidic solution: Problems #1-10

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Problem #8: $\text{Fe} + \text{HCl} \rightarrow$

$\text{HFeCl}_4 + \text{H}_2$. Solution: 1)

This problem poses

interesting problems,

especially with the Cl. The

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key to solving this problem is to eliminate everything not directly involved in the redox. That means the H in HFeCl_4 as well as the Cl in it and HCl. When we do that, this is the unbalanced ...

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Balancing Redox Equations in
Basic Solution Example

Problem Balancing Redox
Equations in Basic Solution
Example Problem by Tyler

DeWitt 4 years ago 10
minutes 103,050 views A

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practice , problem , for how
to balance an , oxidation
reduction , (, redox ,)
reaction in basic , solution
, . The process is similar
to balance

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Solution: Balance the equation using the half-reaction method outlined in the Balance Redox Reaction Example. This reaction is the same one used in the example but was balanced in an acidic environment. The

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example showed the balanced equation in the acidic solution was: $3 \text{ Cu} + 2 \text{ HNO}_3 + 6 \text{ H}^+ \rightarrow 3 \text{ Cu}^{2+} + 2 \text{ NO} + 4 \text{ H}_2\text{O}$ There are six H^+ ions

How to Balance a Redox Reaction in a Basic Solution

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Basic Solution 1. $\text{MnO}_4^- + \text{C}_2\text{O}_4^{2-} \rightarrow \text{MnO}_2 + \text{CO}_2$

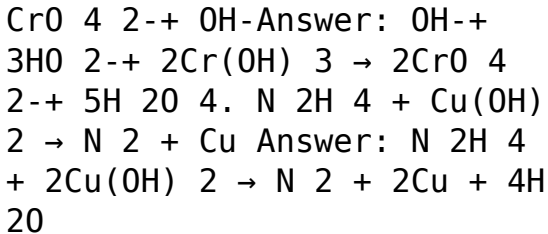
Answer: $4\text{H}^+ + 2\text{MnO}_4^- + 3\text{C}_2\text{O}_4^{2-} \rightarrow 2\text{MnO}_2 + 6\text{CO}_2 +$

8OH^- . $\text{ClO}_4^- + \text{Fe}(\text{OH})_3 \rightarrow \text{Cl}^- + \text{FeO}_4^{2-}$

Answer: $4\text{OH}^- + 3\text{ClO}_4^- + 2\text{Fe}(\text{OH})_3 \rightarrow 3\text{Cl}^- + 2\text{FeO}_4^{2-} +$

$5\text{H}_2\text{O}$. $\text{H}_2\text{O}_2 + \text{Cr}(\text{OH})_3 \rightarrow$

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Oxidation-Reduction Extra

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Practice

The Redox Reaction Class 11 solutions are provided with exercise questions and answers to help them revise the complete syllabus and score good marks in the examinations. You can also

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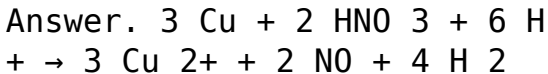
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**NCERT Solutions for Class 11
Chemistry Chapter 8 Redox**

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To summarize: Identify the oxidation and reduction components of the reaction. Separate the reaction into the oxidation half-reaction

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and reduction half-reaction.
Balance each half-reaction
both atomically and
electronically.

Balance Redox Reaction Example Problem

Redox measurement issues.

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The main problems associated with redox measurement is that the probe containing the redox electrode is rather large (about the diameter of a pencil), and when inserted into the sediment, oxygenated water

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from above the sediment
enters around the probe
affecting the redox
potential.

**The complex principles of
redox measurement « Global**

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**Redox reactions questions
(practice) | Khan Academy**

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Some redox titration problems involving manganate(VII) Background

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skills 1. How many moles of manganate(VII) ions, MnO_4^- , are in the following solutions? a) 1000cm^3 of 1M potassium manganate(VII), KMnO_4 . b) 25.0 cm^3 of 1M KMnO_4 . c) 25.0 cm^3 of 0.020M KMnO_4 . Element Ar H

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1.00 N 14.0 0 16.0 S 32.1 K
39.1 Mn 54.9 Fe 55.8 d) 36.5
cm³ of 0.012M ...

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